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(71) Applicant:
AIR PRODUCTS AND CHEMICALS, INC.
Allentown, PA 18195-1501 (US)

(72) Inventors:
• Li, Hong-Xin
Lansdale, PA 19446 (US)

• Emig, Lenore Ann
Whitehall, PA 18052 (US)
• Underwood, Richard Paul
Allentown, PA 18106 (US)

(74) Representative:
Schwabe - Sandmair - Marx
Stuntzstrasse 16
81677 München (DE)

Remarks:

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(54) Method of preparing phosphate catalysts for triethylenediamine production

(57) A new method of making phosphate-based catalysts by mixing phosphoric acid with a substantially water insoluble alkaline earth metal salt such that the phosphorus to alkaline earth metal molar ratio is less than 1. The product, containing alkaline earth metal hydrogen phosphate and the starting alkaline earth metal salt, is filtered and dried. The product can be used as a catalyst in the production of triethylenediamine from, for example, mono- and di-substituted piperazines, such as hydroxyethylpiperazine and aminoethylpiperazine, ethanolamines and substituted ethanolamines, and crude hydroxyethylpiperazine containing piperazine, hydroxyethylpiperazine, bis-hydroxyethylpiperazine, and water.

Description**BACKGROUND OF THE INVENTION**

5 [0001] Organic synthesis by condensation reactions resulting in the loss of a molecule of water or of ammonia is well known in the art. Certain of such reactions are generally effected in the presence of acidic catalysts. An important area in which such acid catalysis has been employed is in cyclization reactions as in the synthesis of triethylenediamine and its C-substituted homologues. The catalysts are typically solid products of the Lewis acid type.

10 [0002] One group of catalysts which have been found to be effective for acid catalyzed organic condensation reactions such as those used to produce triethylenediamine (also referred to herein as TEDA) are phosphate catalysts. The following patents provide examples of such phosphate catalysts.

15 [0003] U. S. 3,297,701 (Brader et al., 1967) discloses the use of metal phosphate catalysts such as aluminum phosphate, calcium phosphate and cobalt phosphate, for synthesis of TEDA and C-substituted TEDA from piperazines or alkanolamines.

20 [0004] U. S. 4,405,784, U. S. 4,514,567, and U. S. 4,521,600 (Wells et al, 1983-1985) disclose the use of strontium monohydrogen phosphate (SrHPO_4), strontium pyrophosphate ($\text{Sr}_2\text{P}_2\text{O}_7$), strontium dihydrogen phosphate ($\text{Sr}(\text{H}_2\text{PO}_4)_2$), the pyrophosphate, monohydrogen phosphate and dihydrogen phosphate of copper, magnesium, calcium, barium, zinc, aluminum, lanthanum, cobalt, nickel, cerium and neodymium, and mixtures thereof, as effective catalysts for organic condensation reactions especially the reaction to form TEDA. The catalysts are made by reacting a soluble salt of one of the metals with the mono- or diphosphate of an alkali metal or ammonium. The pH of the reaction mixture is adjusted to 5 ± 3 in order to precipitate out the mono- or dihydrogen phosphate of the metal.

25 [0005] U. S. 4,757,143 (Vanderpool et al, 1988) discloses the conversion of cyclic and acyclic hydroxyethyl ethylene polyamines to TEDA using a catalyst composed of zirconia or titania to which from 0.5 to about 7 wt % of phosphorous has been thermally chemically bonded in the form of phosphate linkages.

25 [0006] U. S. 5,037,838 (Zimmerman et al, 1991) discloses the conversion of N-hydroxyethyl piperazine to TEDA using a titania-supported tungstopyrophosphate catalyst.

BRIEF SUMMARY OF THE INVENTION

30 [0007] This invention is directed to a new method of making phosphate-based catalysts which are useful for the production of triethylenediamine (TEDA). The catalysts are prepared by mixing phosphoric acid with a substantially water insoluble alkaline earth metal salt such that the phosphorus to alkaline earth metal molar ratio is less than 1. An aqueous slurry of the salt, such as strontium carbonate, calcium carbonate, or barium carbonate, is mixed with an aqueous solution of phosphoric acid to form an alkaline earth hydrogen phosphate. Since less than stoichiometric amounts of alkaline earth metal salt and acid are used, some of the alkaline earth metal salt remains unreacted and is present in the precipitated product. The product, containing alkaline earth metal hydrogen phosphate and the starting alkaline earth metal salt, is filtered and dried. The mixture can be used as a catalyst in the production of TEDA from, for example, mono-and di-substituted piperazines, such as hydroxyethylpiperazine and aminoethylpiperazine, ethanolamines and substituted ethanolamines, and crude hydroxyethylpiperazine containing piperazine, hydroxyethylpiperazine, bis-hydroxyethylpiperazine, and water.

35 [0008] There are several advantages to this method of making the alkaline earth metal hydrogen phosphate compared to the known method of reacting an alkaline earth metal nitrate with sodium hydrogen phosphate, both of which are in solution:

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- 45 • the alkaline earth metal salt does not need to be put into solution before reacting it with phosphoric acid;
- the catalyst does not need to be washed to remove unwanted metals such as sodium; and
- the cost of the raw materials is much lower.

50 It has also been found that the product made by the method of this invention has much better activity and selectivity in making TEDA from hydroxyethylpiperazine than catalysts formed from a known method.

DETAILED DESCRIPTION OF THE INVENTION

55 [0009] Preparation of alkaline earth metal hydrogen phosphates can be carried out by first forming an aqueous slurry of an alkaline earth metal salt which is substantially insoluble in water, such as strontium carbonate, barium carbonate, or calcium carbonate. By "substantially insoluble in water" is meant that the solubility in water at ambient temperature (i.e., approximately 25 °C) is less than 1 part per 100 parts water.

[0010] An aqueous solution of phosphoric acid, for example, an 85 % aqueous solution, is added to the alkaline

earth metal salt slurry, with stirring, in an amount that the molar ratio of phosphoric acid to alkaline earth metal salt is less than 1; preferably less than 0.8. The reaction can be carried out at ambient temperature, i.e., approximately 25 °C, and atmospheric pressure.

[0011] The product precipitate is a mixture of the alkaline earth metal hydrogen phosphate salt and unreacted alkaline earth metal salt. The molar ratio of phosphorous to alkaline earth metal in the product is less than 1; preferably, less than 0.8. The precipitate is filtered and can then be washed, for example with water, although washing is not required. The filtered precipitate can then be dried.

[0012] For use as a catalyst, the product may be employed in the form of irregular particles of a desired size range by breaking up the washed and dried filter cake or in the form of regular shaped pellets obtained by known methods of casting, pelletizing or extruding. The product may also be deposited or otherwise impregnated into the pores of a micro-porous substrate such as alumina, silica, silica-alumina, and the like.

[0013] In using the catalyst of the present invention to catalyze organic condensation reactions, substantially the same conditions may be employed as when using known catalysts. For optimum results, however, some adjustment in temperature, diluent, and/or space rate may be found beneficial.

[0014] A continuous process is typically used in the production of the triethylenediamine. The temperature range is about 285 to 420 °C preferably 300 to 390 °C, the pressure range is about 0.1 to 1.5 atmospheres (101.4 to 152.03 kPa), preferably 0.3 to 1.0 atm (30.3 to 101.4 pKa), and the liquid hourly space velocity (LHSV) of organic feedstock per volume of catalyst is in the range of about 0.05 to 1.5, preferably 0.1 to 0.3.

[0015] The reaction can be carried out in the presence of an inert gas such as nitrogen, argon or helium.

[0016] In the preparation of TEDA, the organic feedstock includes mono- and di-substituted piperazines selected from the group consisting of hydroxyethylpiperazine, aminoethylpiperazine, ethanolamines, and substituted ethanolamines. The catalysts of this invention are relatively unaffected by the purity of the feedstock. For example, high conversion and good yields can be obtained from crude hydroxyethylpiperazine containing, in addition to the hydroxyethylpiperazine, piperazine, bis-hydroxyethylpiperazine, and water.

[0017] The invention will be further clarified by a consideration of the following examples, which are intended to be purely exemplary of the invention.

EXAMPLE 1

30 PREPARATION OF CATALYST

[0018] A slurry was prepared by adding 30.1 g of SrCO_3 (supplied by CPC) to 40.02 g of de-ionized water. To this slurry, 11.6 g of 85% phosphoric acid was added with stirring. Carbon dioxide was released during the reaction. The solid was filtered, washed with de-ionized water, and dried at 110 °C. The solid contained a mixture of SrHPO_4 and SrCO_3 . Chemical analysis showed that the solid had a P/Sr ratio of 0.54. The BET surface area was $6 \text{ m}^2/\text{g}$.

EXAMPLE 2

40 PREPARATION OF TEDA FROM HYDROXYETHYL PIPERAZINE USING A CATALYST PREPARED FROM STRONIUM CARBONATE

[0019] The reaction was carried out in a continuous flow, tubular reactor under atmospheric pressure. The catalyst of Example 1, pelletized to 18-35 mesh particles, was loaded into the reactor and heated to 340 °C under dry nitrogen flow. An aqueous solution of 25 % hydroxyethyl piperazine (HEP) was pumped into the reactor at 1.5 ml/min. Nitrogen was co-fed into the reactor at 22 ml/min. Reaction effluent was collected and analyzed by gas chromatograph. Results are shown in the table below.

EXAMPLE 3

50 PREPARATION OF TEDA FROM HYDROXYETHYL PIPERAZINE USING A CATALYST PREPARED FROM STRONIUM CARBONATE

[0020] The catalyst was prepared as in Example 1 except that the SrCO_3 was supplied by Aldrich. TEDA was prepared according to the procedure of Example 2. Results are shown in the table below.

EXAMPLE 4

PREPARATION OF TEDA FROM HYDROXYETHYL PIPERAZINE USING A CATALYST PREPARED FROM CALCIUM CARBONATE

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[0021] The catalyst was prepared according to the procedure of Example 1 except that calcium carbonate was used instead of strontium carbonate. TEDA was prepared according to the procedure of Example 2. Results are shown in the table below.

10 COMPARATIVE EXAMPLE 5

PREPARATION OF STRONTIUM HYDROGEN PHOSPHATE BY A PRIOR ART PROCESS AND USE OF THE CATALYST IN TEDA SYNTHESIS

15 **[0022]** A strontium hydrogen phosphate catalyst was prepared according to the procedure of U.S. 4,405,784 in which a solution of strontium nitrate ($\text{Sr}(\text{NO}_3)_2$) was reacted with sodium hydrogen phosphate (Na_2HPO_4) and the pH of the reaction mixture was adjusted to about 5.5. The product was removed by filtration, washed with water, and dried. The catalyst was used in the production of TEDA according to the procedure of Example 2. Results are shown in the table below.

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25	Ex.	Catalyst	Reactants for Forming Catalyst	Hours on Stream	% HEP Converted	Products		
						TEDA wt %	Piperazine wt %	Other wt %
30	2	$\text{SrHPO}_4/\text{SrCO}_3$ P/Sr=0.5 BET=6 m^2/g	SrCO_3 (CPC) + 85% H_3PO_4	30	80.3	82.7	4.7	12.6
35	3	$\text{SrHPO}_4/\text{SrCO}_3$ P/Sr=0.5 BET=6.9 m^2/g	SrCO_3 (Aldrich) + 85% H_3PO_4	50	84.9	85.5	4.5	10.0
40	4	$\text{CaHPO}_4/\text{CaCO}_3$ P/Sr=0.5 BET=17 m^2/g	CaCO_3 (Aldrich) + 85% H_3PO_4	30	82.7	83.4	7.4	9.3
45	5	SrHPO_4 P/Sr=1 BET=9 m^2/g	$\text{Sr}(\text{NO}_3)_2 + \text{Na}_2\text{HPO}_4$	48	66.7	73.1	4.8	22.1

50 Reaction conditions: 1.0 cc catalyst (18-35 mesh); temperature = 340 °C; pressure = 1 atm; feed = 1.5 ml/h of 25 wt % HEP; nitrogen flow = 22 cc/min.

[0023] Data for examples 2, 3, and 4, in which a mixture of alkaline earth metal hydrogen phosphate and carbonate were used, compared to example 5, in which pure strontium hydrogen phosphate was used, show that use of the catalyst of this invention unexpectedly resulted in higher conversion of hydroxyethylpiperazine to TEDA and higher selectivity to TEDA. In addition, the higher conversion and higher selectivity of the catalyst of this invention was achieved with lower amounts of the strontium or calcium hydrogen phosphate. The data of examples 2, 3 and 4 also show that improved conversion and selectivity were obtained over a wide range of catalyst surface area; i.e., 6 m^2/g to 17 m^2/g .

Claims

1. A method for preparing a catalyst comprising an alkaline earth metal hydrogen phosphate and an alkaline earth metal salt which comprises

5 combining an aqueous phosphoric acid solution with an aqueous slurry of a substantially water insoluble alkaline earth metal salt such that the molar ratio of phosphoric acid to alkaline earth metal salt is less than 1, to form a precipitate of the alkaline earth metal hydrogen phosphate containing residual alkaline earth metal salt; filtering the precipitate of the alkaline earth metal hydrogen phosphate containing residual alkaline earth metal
10 salt; and
 drying the precipitate.

15 2. The method of claim 1 wherein the alkaline earth metal salt is selected from the group consisting of strontium carbonate, calcium carbonate, and barium carbonate.

16 3. The method of claim 2 wherein the alkaline earth metal salt is strontium carbonate.

17 4. The method of claim 2 wherein the alkaline earth metal salt is calcium carbonate.

20 5. The method of claim 2 wherein the molar ratio of phosphoric acid to alkaline earth metal salt is less than 0.8.

25 6. A method for preparing triethylenediamine comprising
 contacting a compound selected from the group consisting of hydroxyethylpiperazine, crude hydroxyethylpiperazine, N-aminoethyl piperazine, ethanolamines, and substituted ethanolamines, with a catalyst comprising a combination of an alkaline earth metal hydrogen phosphate and an alkaline earth metal salt wherein the molar ratio of phosphorus to alkaline earth metal in the combination is less than 1.

30 7. The method of claim 6 wherein the molar ratio of phosphorus to alkaline earth metal in the combination is less than 0.8.

31 8. The method of claim 7 wherein the alkaline earth metal is strontium, calcium, or barium, and the salt is carbonate.

32 9. The method of claim 8 wherein the compound is hydroxyethylpiperazine or crude hydroxyethylpiperazine.

35 10. The method of claim 8 wherein the compound is crude hydroxyethylpiperazine.

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EUROPEAN SEARCH REPORT

Application Number

EP 00 11 0213

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
A	EP 0 644 156 A (FUJI CHEM IND CO LTD) 22 March 1995 (1995-03-22) ----		B01J27/18 C01B25/32 C07D487/08
A	EP 0 199 912 A (BENCKISER KNAPSACK GMBH) 5 November 1986 (1986-11-05) ----		
D, A	EP 0 069 322 A (AIR PROD & CHEM) 12 January 1983 (1983-01-12) -----		
TECHNICAL FIELDS SEARCHED (Int.Cl.7)			
B01J C01B C07D			
The present search report has been drawn up for all claims			
Place of search	Date of completion of the search	Examiner	
THE HAGUE	25 August 2000	Thion, M	
CATEGORY OF CITED DOCUMENTS			
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			
T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document			

ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.

EP 00 11 0213

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

25-08-2000

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
EP 0644156	A	22-03-1995	JP 2700141 B JP 7118005 A CA 2127740 A DE 69407706 D DE 69407706 T US 5486365 A	19-01-1998 09-05-1995 18-03-1995 12-02-1998 16-04-1998 23-01-1996
EP 0199912	A	05-11-1986	DE 3515695 A AT 49950 T DE 3668608 D US 4755367 A	06-11-1986 15-02-1990 08-03-1990 05-07-1988
EP 0069322	A	12-01-1983	US 4405784 A US 4521600 A US 4501889 A US 4446320 A AT 36654 T AU 8465682 A BR 8203774 A CA 1198430 A DE 3278930 D ES 513523 D ES 8401434 A JP 1013701 B JP 1526468 C JP 58017839 A MX 170969 B US 4514567 A ZA 8204133 A CA 1198428 A CA 1198429 A US 4582904 A EP 0111928 A JP 1606897 C JP 2022756 B JP 59152350 A	20-09-1983 04-06-1985 26-02-1985 01-05-1984 15-09-1988 06-01-1983 21-06-1983 24-12-1985 29-09-1988 16-12-1983 01-03-1984 07-03-1989 30-10-1989 02-02-1983 22-09-1993 30-04-1985 27-04-1983 24-12-1985 24-12-1985 15-04-1986 27-06-1984 13-06-1991 21-05-1990 31-08-1984